

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium

KEY



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



- (A) 95²
 (B) 95
 (C) (95)²
 (D) $\frac{1}{95}$
 (E) $\frac{1}{(95)^2}$



$C_2H_2(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
 II. Changes in the particle size of the platinum catalyst
 III. Changes in the temperature of the reaction system

- (A) III only
 (B) I and II only
 (C) I and III only
 (D) II and III only
 (E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

- (A) Removing a small amount of $O_2(g)$ *R*
 (B) Removing a small amount of $BaO(s)$ *X*
 (C) Adding He gas to the vessel *X*
 (D) Lowering the temperature *L*

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? *increase P*

- (A) $SrCO_3(s) \rightleftharpoons SrO(s) + CO_2(g)$
 (B) $Cl_2 + COCl_2(g) \rightleftharpoons 2 COCl(g)$
 (C) $CO(g) + H_2O(g) \rightleftharpoons CO_2(g) + H_2(g)$
 (D) $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$
 (E) $2 NO_2(g) \rightleftharpoons N_2O_4(g)$
- shift towards fewer & make gas*



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

- (A) 3
 (B) 6
 (C) 8
 (D) 16
 (E) 36
- $X + Y \rightleftharpoons 2Z$
 $4M \quad 4M \quad 0$
 $-x \quad -x \quad +2x$
 $1M \quad 1M \quad 6M$
 $2x = 6M \Rightarrow x = 3M$
 $K_c = \frac{[Z]^2}{[X][Y]} = \frac{6^2}{1 \times 1} = 36$

6. The synthesis of $CH_3OH(l)$ from $CO(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease Kc

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

- (A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.
 (B) K_c will increase because there are more reactant molecules than product molecules.
 (C) K_c will decrease because the reaction is exothermic.
 (D) K_c is constant and will not change.

2002 Ap Chemistry Multiple Choice Answers Explained

Gerald Litwack



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Tools of Radio Astronomy T. L. Wilson, Kristen Rohlf, Susanne Hüttemeister, 2008-11-14 Four significant factors have led us to update this text. The first is the breathtaking progress in technology especially in receiver and digital techniques. The second is the advance of radio astronomy to shorter wavelengths and the increased availability of astronomical satellites. The third is a need to reorganize some of the chapters in order to separate the basic theory that seldom changes from practical aspects that change often. Finally it is our desire to enhance the text by including problem sets for each chapter. In view of this ambitious plan we have expanded the number of authors. In the reorganization of this edition we have divided Chap 4 of the 4th edition into two Chaps 4 and 5. The first remains Chap 4 with a slightly different title Signal Processing and Receivers Theory. This was expanded to include digital processing and components including samplers and digitizers. In Chap 5 Practical Receiver Systems we have relegated the presentations of maser and parametric amplifier front ends which are no longer commonly used as microwave receivers in radio astronomy to a short section on historical developments and We have retained and improved the presentations of current state of the art devices cooled transistor and superconducting front ends. We have also included descriptions of local oscillators and phase lock loops. Chapters 5 and 6 in the 4th edition has now become Chap 6 Fundamentals of Antenna Theory and Chap 7 **Physics of Comets** K. S. Krishna Swamy, 2010 This revised edition places a unique emphasis on all the new results from ground based satellites and space missions detection of molecule H₂ and prompt emission lines of OH for the first time discovery of X rays in comets observed diversity in chemical composition among comets the puzzle of the constancy of spin temperature the well established mineralogy of cometary dust extensive theoretical modeling carried out for understanding the observed effects the similarity in the mineralogy of dust in circumstellar shell of stars comets meteorites asteroids and IDPs thus indicating the generic relationship between them

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