

18

REACTION RATES AND EQUILIBRIUM

SECTION 18.1 RATES OF REACTION (pages 541–547)

This section explains what is meant by the rate of a chemical reaction. It also uses collision theory to describe the rate of a chemical reaction as determined by the reaction conditions.

Collision Theory (pages 541–544)

1. How can rates of chemical change expressed?

Rates of chemical change are usually expressed in dimensions of moles/liters/second per unit time.

2. Look at Figure 18.1 on page 541. In a typical reaction, as time passes, the amount of reactants decreases and the amount of products increases.

3. What does collision theory say about the orientation of atoms, ions, or molecules reacting to form products when they collide?

Collision theory states that the molecules must have enough kinetic energy and the correct orientation to form products.

4. Look at the space below. One shows a collision that results in the formation of product. Label it effective collision. Label the other collision's distribution of atoms.



5. In the following sequence from an Ideal Particle lacking the necessary kinetic energy to react because it is not exchanged when they collide. atom

6. Look at Figure 18.5 on page 545. Which arrangement of atoms contains the least amount of energy?

- a. reactants
- b. activated complex
- c. products**

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Chapter 18 Reaction Rates And Equilibrium Workbook

M Mosston

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