

17

THERMOCHEMISTRY

Reaction that releases heat is exothermic
 Reaction that absorbs heat is endothermic
 Exothermic: $\Delta H < 0$
 Endothermic: $\Delta H > 0$

Practice Problems

1. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK

1. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

2. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

3. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

4. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

5. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

6. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

7. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

8. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

9. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

10. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

SECTION 17.2 MEASURING AND EXPRESSING ENTHALPY CHANGES

1. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

2. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

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4. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

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6. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

7. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

8. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

SECTION 17.3 HEAT IN CHANGES OF STATE

Chapter 17 Thermochemistry Practice Problems

Elton N. Kaufmann



Chapter 17 Thermochemistry Practice Problems:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific, detailed, step-by-step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step-by-step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid-base reactions, which often persists through working with acid-base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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Computational Thermochemistry Karl K. Irikura, 1998. Comprises 20 contributions which grew from the August 1996 symposium. Representative paper topics include estimating phase change enthalpies and entropies, electrostatic covalent model.

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