

# % Concentration

## Example Problem 1 (Solid in a Liquid)

What is the concentration in %w/v of a solution containing 39.2 g of potassium nitrate in 177 mL of solution?

$$\% \text{ (w/v)} = \frac{\text{mass solute}}{\text{volume solution}} \times 100 = \frac{39.2 \text{ g}}{177 \text{ mL}} \times 100 = 22.1 \% \text{ w/v}$$

## Example Problem 2 (Liquid in a Liquid)

What is the concentration in %v/v of a solution containing 3.2 L of ethanol in 6.5 L of solution?

$$\% \text{ (v/v)} = \frac{\text{volume solute}}{\text{volume solution}} \times 100 = \frac{3.2 \text{ L}}{6.5 \text{ L}} \times 100 = 49 \% \text{ v/v}$$

# Chapter 15 Section 152 Solution Concentration Answers

**J Rink**



## **Chapter 15 Section 152 Solution Concentration Answers:**

**Chemistry Workbook For Dummies** Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence

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