

Study Packet: Stoichiometry

Stoy - kee - ahm - eh - tree

Set I. Conversion Factors

Consider the following reaction:



a) **2 mole $\text{H}_3\text{PO}_4 = 6 \text{ mole } \text{H}_2\text{O}$**

b) **3 moles $\text{Hg}(\text{OH})_2 = 1 \text{ mole } \text{Hg}_3(\text{PO}_4)_2$**

c) **3 moles $\text{Hg}(\text{OH})_2 = 2 \text{ moles } \text{H}_3\text{PO}_4$**

d) **1 mole $\text{Hg}_3(\text{PO}_4)_2 = 6 \text{ moles } \text{H}_2\text{O}$**

What is the mole to mole relationship for phosphoric acid to water?

What is the mole to mole relationship for $\text{Hg}(\text{OH})_2$ to $\text{Hg}_3(\text{PO}_4)_2$?

What is the mole to mole relationship for $\text{Hg}(\text{OH})_2$ to H_3PO_4 ?

What is the mole to mole relationship for $\text{Hg}_3(\text{PO}_4)_2$ to H_2O ?

Balance the following equations	Write out all of the possible mole ratios
2. <u>2</u> $\text{SO}_2 + \underline{\quad} \text{O}_2 \rightarrow \underline{2}$ SO_3	2. 2 mol $\text{SO}_2 = 1 \text{ mole } \text{O}_2$, 2 moles $\text{SO}_2 = 2 \text{ moles } \text{SO}_3$, 1 mole $\text{O}_2 = 2 \text{ mole } \text{SO}_3$
3. <u>1</u> $\text{PCl}_3 + \underline{\quad} \text{Cl}_2 \rightarrow \underline{1}$ PCl_5	3. Everything is 1:1
4. <u>4</u> $\text{NH}_3 + \underline{3}$ $\text{O}_2 \rightarrow \underline{2}$ $\text{N}_2 + \underline{6}$ H_2O	4. 4 mol $\text{NH}_3 = 3 \text{ mol } \text{O}_2$, 4 mol $\text{NH}_3 = 2 \text{ mol } \text{N}_2$, 4 mol $\text{NH}_3 = 6 \text{ mol } \text{H}_2\text{O}$, 3 mol $\text{O}_2 = 2 \text{ mol } \text{N}_2$, 3 mol $\text{O}_2 = 6 \text{ mol } \text{H}_2\text{O}$, 2 mol $\text{N}_2 = 6 \text{ mol } \text{H}_2\text{O}$
5. <u>1</u> $\text{Fe}_2\text{O}_3 + \underline{3}$ $\text{CO} \rightarrow \underline{2}$ $\text{Fe} + \underline{3}$ CO_2	5. 1 mol $\text{Fe}_2\text{O}_3 = 3 \text{ mol } \text{CO}$, 1 mol $\text{Fe}_2\text{O}_3 = 2 \text{ mol } \text{Fe}$, 1 mol $\text{Fe}_2\text{O}_3 = 3 \text{ mol } \text{CO}_2$, 3 mol $\text{CO} = 2 \text{ mol } \text{Fe}$, 3 mol $\text{CO} = 3 \text{ mol } \text{CO}_2$, 2 mol $\text{Fe} = 3 \text{ mol } \text{CO}_2$

Set II. Mole to mole and mole to mass

On a separate sheet of paper balance the equations for each problem and solve the questions. Don't forget to make sure the reaction is balanced!



1. "sulfur burns in oxygen gas to form sulfur dioxide"

a. **1.56 mol S**

How many moles of sulfur must be burned to give 100.0 g of SO_2 ?

b. **10 mol S**

How many moles of sulfur must be burned to completely react with 10 moles of oxygen?

c. **1.93×10^{25} mc O_2**

If there are 32 moles of oxygen, how many molecules of oxygen are there?



a. **2.56×10^7 mg $\text{Ba}(\text{NO}_3)_2$**

How many mg of $\text{Ba}(\text{NO}_3)_2$ can be formed from 196. moles of HNO_3 ?

b. **14.9 mol HNO_3**

If 7.45 moles of water are formed during this reaction, how many moles of HNO_3 were used?



a. **9.85 mol AgNO_3**

How many moles of AgNO_3 are needed to produce 9.85 mol of silver.

b. **15.5 mole Cu**

How many moles of copper produce 5.00 kg of silver?

Chapter 12 Stoichiometry Packet

Puja Mehta



Chapter 12 Stoichiometry Packet:

Chapter 12 Stoichiometry Packet Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chapter 12 Stoichiometry Packet**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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