

Limiting Reactants

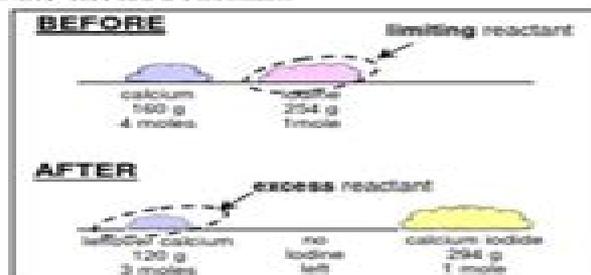
Chem Worksheet 12-3

Name _____

When an automobile production plant runs out of tires no more cars can be produced even if there are still plenty of engines, bodies, seats, and other parts. When one of the reactants in a chemical reaction runs out the reaction stops as well. The substance that limits production in a chemical reaction is known as the **limiting reactant**. The reactant that is left over when the reaction stops is called the **excess reactant**.



Consider the chemical reaction between calcium and iodine. When 160 g (4 moles) of calcium reacts with 254 g (1 mole) of iodine the reaction makes 294 g (1 mole) of calcium iodide. Although there is enough calcium to make more calcium iodide, the iodine runs out first. Since all of the iodine gets used up it is called the limiting reactant. According to the balanced equation, if one mole of iodine reacts, one mole of calcium will react. This means that there are still 3 moles of calcium left. Because calcium is left over it is called the excess reactant.



Example

What mass of iron (II) sulfide will be produced if 9.68 g of iron reacts with 6.28 g of sulfur? $__ \text{Fe} + __ \text{S} \rightarrow __ \text{FeS}$

- balance the equation



- perform a calculation for each reactant:

$$\text{Fe: } \frac{9.68 \text{ g Fe}}{1} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} \times \frac{1 \text{ mol FeS}}{1 \text{ mol Fe}} \times \frac{87.91 \text{ g FeS}}{1 \text{ mol FeS}} = 15.24 \text{ g FeS}$$

$$\text{S: } \frac{6.28 \text{ g S}}{1} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} \times \frac{1 \text{ mol FeS}}{1 \text{ mol S}} \times \frac{87.91 \text{ g FeS}}{1 \text{ mol FeS}} = 17.22 \text{ g FeS}$$

- whichever reactant makes less product is the limiting reactant:

Iron is the limiting reactant. There is enough sulfur to make 17.22 g FeS, but only enough iron to make 15.24 g of FeS.

Find the amount of product formed in each of the following reactions.

- What mass of H_2O will be produced if 9.5 g of H_2 reacts with 1.2 g of O_2 ?
 $__ \text{H}_2 + __ \text{O}_2 \rightarrow __ \text{H}_2\text{O}$
- If 1.85 g of $\text{Mg}(\text{OH})_2$ reacts with 3.71 g of HCl , how much MgCl_2 is produced? What is the limiting reactant?
 $__ \text{Mg}(\text{OH})_2 + __ \text{HCl} \rightarrow __ \text{MgCl}_2 + __ \text{H}_2\text{O}$
- What mass of AgCl is produced when 53.42 g of AgNO_3 reacts with 14.19 g of NaCl ?
 $__ \text{AgNO}_3 + __ \text{NaCl} \rightarrow __ \text{AgCl} + __ \text{NaNO}_3$
- If 14.7 g of calcium is placed in 11.5 g of water, what mass of hydrogen gas is produced?
 $__ \text{Ca} + __ \text{H}_2\text{O} \rightarrow __ \text{Ca}(\text{OH})_2 + __ \text{H}_2$
- What mass of potassium hydroxide is formed when 8.2 g of potassium oxide is added to 1.3 g of water?
 $__ \text{K}_2\text{O} + __ \text{H}_2\text{O} \rightarrow __ \text{KOH}$
- What mass of aluminum chloride could be made from 8.1 g of aluminum and 4.2 L of chlorine at STP?
 $__ \text{Al} + __ \text{Cl}_2 \rightarrow __ \text{AlCl}_3$
- If 5.26 L of nitrogen monoxide and 7.64 L of oxygen are combined, what mass of nitrogen dioxide is formed? Assume conditions are STP.
 $__ \text{NO} + __ \text{O}_2 \rightarrow __ \text{NO}_2$
- If 18.1 g of silicon tetrachloride reacts with 8.4 L of hydrogen at STP, what mass of silicon is formed?
 $__ \text{SiCl}_4 + __ \text{H}_2 \rightarrow __ \text{Si} + __ \text{HCl}$
- What mass of $\text{Cu}(\text{OH})_2$ is produced in the reaction of 7.6 g $\text{Cu}(\text{NO}_3)_2$ with 6.2 g of KOH ?
 $__ \text{Cu}(\text{NO}_3)_2 + __ \text{KOH} \rightarrow __ \text{Cu}(\text{OH})_2 + __ \text{KNO}_3$
- What mass of carbon dioxide is formed when 64 kg of ethylene is burned in 142 kg of oxygen?
 $__ \text{C}_2\text{H}_4 + __ \text{O}_2 \rightarrow __ \text{CO}_2 + __ \text{H}_2\text{O}$

Chapter 12 Section 12 3 Limiting Reactants Answers

Morris Hein, Susan Arena



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