

Chemistry Study Guide

Chapter 1

- **1.1: Chemistry-The Science of Everyday Experience**
 - Chemistry = The study of matter- its composition, properties, and transformations
 - Matter = Anything that has mass and takes up volume
- **1.2: States of Matter**
 - Solid: Definite volume, Definite shape. Particles are close together and are in a regular 3-dimensional array
 - Liquid: Definite volume, Indefinite shape. Particles are close together but randomly move around.
 - Gas: Indefinite volume, Indefinite shape. Particles far apart.
- Physical properties are those that can be observed or measured without changing the composition of the material
 - Ex: Boiling point, melting point, solubility, color, odor
- A physical change alters a substance without changing its composition
 - Ex: Changes in state, melting point, boiling point, solubility, color, odor.
- Chemical properties are those that determine how a substance can be converted to another substance
- A chemical change, or a chemical reaction, converts one material to another
 - Ex: Conversion of Hydrogen and Oxygen to water

1.3: Classification of Matter

- Pure Substance: Composed of a single component and has a constant composition, regardless of the sample size and the origin of the sample
 - Can't be broken down to other pure substances by any physical change
 - Ex: Water, Table sugar
 - A pure substance can be classified as either an element or compound
 - Element: Pure substance that cannot be broken down into simpler substances by a chemical reaction
 - 1. Ex: Nitrogen, Hydrogen
 - Compound: Pure substance formed by chemically combining 2 or more elements
 - 1. Ex: Water
 - ii. Mixture: Composed of more than one substance
 - The composition of a mixture can be separated into its components by physical changes.
 - Ex: Sugar+Water
 - iii. Mass: Measure of the amount of matter in an object
 - iv. Weight: Force that matter feels due to gravity
- **1.5: Sig Figs**
 - All non-zero digits are always significant
 - Rules to Determine When a Zero Is A Sig Fig

Chapter 12 Chemistry Study Guide

John A. Olmsted, Gregory M. Williams



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