

## Chapter Test B

### Chapter: Stoichiometry

**PART I** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Knowing the mole ratio of a reactant and product in a chemical reaction would allow you to determine
- the energy released in the reaction.
  - the speed of the reaction.
  - the mass of the product produced from a known mass of reactant.
  - whether the reaction is reversible.
- \_\_\_\_\_ 2. In the equation  $6\text{CO}_2 + 6\text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2$ , the mole ratio of water to oxygen is
- 2:1.
  - 1:1.
  - 3:2.
  - 6:8.
- \_\_\_\_\_ 3. In the equation  $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$ , the mole ratio of calcium hydroxide to aluminum hydroxide is
- 1:3.
  - 2:3.
  - 1:1.
  - 3:2.
- \_\_\_\_\_ 4. For the reaction represented by the equation  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ , how many moles of nitrogen are required to produce 18 mol of ammonia?
- 9
  - 18
  - 27
  - 36
- \_\_\_\_\_ 5. In the reaction represented by the equation  $\text{A} + \text{B} \rightarrow \text{C} + \text{D}$ , if there is more reactant B than is required to completely react with all of A, then
- A is the limiting reactant.
  - B is the limiting reactant.
  - there is no limiting reactant.
  - no product can be formed.

# Chapter 11 Stoichiometry Test

**Mary Millhollon, Richard H. Langley**



## Chapter 11 Stoichiometry Test:

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