

Assessment

Chapter Test B - For Practice!

Chapter: Atoms: The Building Blocks of Matter

PART I On the line at the left of each statement, write the letter of the choice that best completes the statement or best answers the question.

- C 1. The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays are composed of
a. energy.
b. positive particles.
 c. negative particles.
d. neutral particles.
- C 2. The basic principles of atomic theory were first conceived by
a. Avogadro.
b. Bohr.
 c. Dalton.
d. Rutherford.
- C 3. The law of multiple proportions is demonstrated by the compounds
a. FeCl_2 and $\text{Fe}(\text{SO}_4)_2$.
b. O_2 and O_3 .
 c. CO and CO_2 .
d. FeCl_2 and $\text{Fe}(\text{NO}_3)_2$.
- B 4. Neutral atoms of the same element can differ in
a. chemical properties.
 b. mass number.
c. atomic number.
d. number of protons and electrons.
- D 5. Dalton's atomic theory helped to explain the law of conservation of mass because it stated that atoms
a. cannot combine.
b. are indivisible.
c. all have the same mass.
 d. cannot be created or destroyed.
- C 6. Millikan's experiments led to the determination
a. that the electron carries no charge.
b. of the electron's positive charge.
 c. of the approximate value of the electron's mass.
d. that the electron has no mass.

Chapter 1biology Test Answers

George Frederick Miller



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