

Unit 7 Acids and Bases Study Guide

Chemistry Standard Set 5

Chapter 19

Key Vocabulary Terms

- | | | |
|------------------------|-------------------------|----------------------|
| 1. Arrhenius Acid | 2. Arrhenius Base | 3. Hydrogen Donor |
| 4. Bronsted-Lowry Acid | 5. Bronsted-Lowry Base | 6. Hydrogen Acceptor |
| 7. Dissociation | 8. Partial Dissociation | 9. Hydronium |
| 10. Neutrality | 11. Buffer | 12. Amphoteric |
| 13. pH Scale | 14. Logarithmic Scale | |

Concepts

1. Observable properties of acids, bases, and salt solutions
2. Distinguish between the Arrhenius and the Bronsted-Lowry definitions of acid or bases.
3. An acid is a substance that, when dissolved in water, forms hydrogen ions (or protons, H^+). A base is a substance that, when dissolved in water, forms hydroxide ions (OH^-).
4. Distinguish between strong acids and bases that fully dissociate and weak acids and bases that partially dissociate.
5. Problems and sources of acid rain.
6. How antacids work in the body.

Items for Memorization

Litmus paper show that litmus is red in the presence of an acid and blue with a base. Phenolphthalein is colorless with an acid and pink with a base.

pH of pure water is 7 and amphoteric.

$pH + pOH = 14.000$; As the pH increases, the concentration of H^+ decreases.

Acids

- sour taste
- react with bases to form salt and water (neutralization)
- pH less than 7
- changes blue litmus paper to red
- changes bicarbonate indicator from red to yellow
- react with reactive metals (like iron, zinc, and aluminum) to produce hydrogen gas
- reacts with limestone ($CaCO_3$) to produce carbon dioxide.

Bases

- Taste bitter (like soap)
- feel slippery to the touch
- react with acids to form salt and water (neutralization)
- change red litmus paper to blue
- pH more than 7
- if soluble in water, known as alkali

Arrhenius concept. An acid is a substance that, when dissolved in water, forms hydrogen ions (or protons, H^+). A base is a substance that, when dissolved in water, forms hydroxide ions (OH^-). **Bronsted-Lowry concept.** An acid is a proton donor; a base is a proton acceptor.

Skills

1. Determine whether a substance is an acid or base based on its pH.
2. Identify acids and bases encountered in and around the home (baking soda, drain cleaner, vinegar, citrus fruits, battery acids, antacids, household ammonia)
3. Write simple equations to show how acids dissociate to release H^+ in solutions.
4. Ability to calculate pH, or pOH based on the hydrogen ion concentration.

Chapter 19 Chemistry Study Guide Acids Bases

David R. Klein, Laurie S. Starkey



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