

CHAPTER 15 REVIEW

Chemical Equilibrium

Teacher Notes and Answers

Question 10

Equilibrium 1

Reaction Equation



- (a) constant
- (b) constant

(c) The values of the forward and reverse rates are affected by increasing the temperature. The forward rate increases and the reverse rate decreases.

- (d) (i) 1
- (ii) 2

Question 11

Reaction Equation



- (a) constant
- (b) constant
- (c) constant
- (d) constant

(e) These results are not consistent with equilibrium. Equilibrium constants are temperature-dependent. The temperature is not constant here.

(f) The reaction is exothermic. An increase in temperature will shift the equilibrium to the left, favouring the reactants.

(g) The reaction is exothermic. An increase in temperature will shift the equilibrium to the left, favouring the reactants.

(h) The reaction is exothermic. An increase in temperature will shift the equilibrium to the left, favouring the reactants.

Question 12

Reaction Equation



(a) The reaction is exothermic. An increase in temperature will shift the equilibrium to the left, favouring the reactants.



(b) The reaction is exothermic. An increase in temperature will shift the equilibrium to the left, favouring the reactants.



(c) The reaction is exothermic. An increase in temperature will shift the equilibrium to the left, favouring the reactants.

Chapter 18 Section Review Answers

**Marcus Tullius Hun, New York (State).
Supreme Court**

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