

Chapter 15 and 16: Water and Solutions

Section 1: Properties of Solutions

- Define the following:
 - Solute **the substance being dissolved**
 - Solvent **the substance doing the dissolving**
 - Solvation **the process of dissolving**
- What are the four factors that affect solution formation? How and why do they have an effect?
 - Composition of solute and solvent – polarity and attraction**
 - Stirring – agitating the particles for more contact**
 - Temperature – increased temp = increased kinetic energy = increased collisions = dissolve**
 - Surface Area – increased surface area means more area for interaction/collisions**
- Dissolving involves an exchange process between two processes: **Solvation and Crystallization**
- What are the three types of solutions and how are they different?
 - Unsaturated – less than maximum amount of solute**
 - Saturated – maximum amount of solute**
 - Supersaturated – more than maximum amount of solute**
- Oil and water are examples of two miscible **immiscible** liquids. (Circle one)
- Temperature **increases** the solubility of solids and **decreases** the solubility of gases.
- How does pressure affect the solubility of solids? Of gases? **No effect on solids, increase pressure to increase gas solubility**
- What is Henry's law? **$S_1 / P_1 = S_2 / P_2$**
- The solubility of a gas at 12 atm is 1.23 g/L. What is the solubility of this gas at 5 atm? **0.513 g/L**
- The solubility of a gas at 740 mmHg is 0.97 g/L. What pressure produces a solubility of 2.1 g/L? **1602 mmHg**
- What is the general solubility rule? **Like dissolves like**

Chapter 15 Review Answer Key

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