

Unit 8: ACIDS AND BASES

Chapter 15: Acids and Bases

15.1: What are Acids and Bases?

Physical and Chemical Properties of Acid and Base

Acids	Bases
Taste Sour (Citric Acids).	Taste Bitter.
Burning Sensation (Stomach Acid).	Feels Slippery (Detergent, Degreaser).
Corrosive with Metals (reacts to give off $H_2(g)$).	Alkaline in Nature (NaOH, Baking Soda).
Conducts Electricity (Strong or Weak electrolytes)	Conducts Electricity (Strong or Weak electrolytes)
Red litmus remains Red; Blue litmus turns Red.	Red litmus turns Blue; Blue litmus remains Blue.
Bromothymol Blue turns Yellow	Bromothymol Blue turns Blue.
Phenolphthalein turns Colourless.	Phenolphthalein turns Pink.
pH < 7	pH > 7



Conceptual Definition: - an explanation that attempts to describe why things are the way they are.

Arrhenius Concept: - acids are H^+ (proton) producers and bases are OH^- producers.



Bronsted-Lowry Model: - acids and bases **react with water** to dissociate where acids are H^+ (proton) donors and bases are H^+ (proton) acceptors.
- first proposed by Johannes Bronsted and Thomas Lowry.

Hydronium Ion: - an ion formed when an acid "donated" H^+ ion combined with a H_2O molecule to form a H_3O^+ ion (hydronium ion).
- essentially has the same function as a H^+ ion, but H_3O^+ denotes that we are using the Bronsted-Lowry model.



Chapter 15 Acids Bases Chapter Review

L Cohen



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