

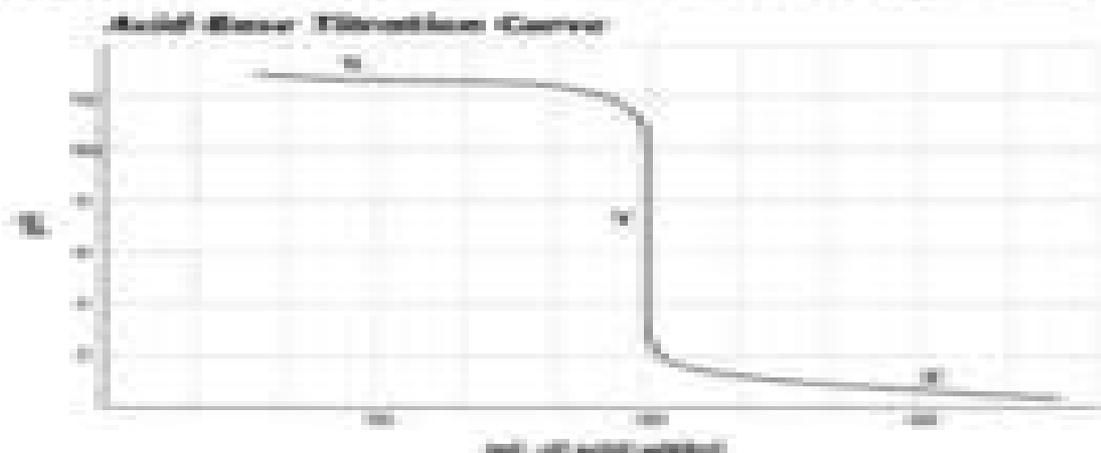
## CHAPTER 14 REVIEW

### Acid-Base Titration and pH

#### SECTION 2

**SHORT ANSWER** Answer the following questions in the space provided.

1. Below is a pH curve from an acid-base titration. The x-axis labels show points A, B, C, and D.



- \_\_\_\_\_ a. Which point represents the equivalence point?
- \_\_\_\_\_ b. At which point is there excess acid in the system?
- \_\_\_\_\_ c. At which point is there excess base in the system?
- \_\_\_\_\_ d. If the three solutions in (b), (c), and (d) were to some equivalence of 100 mL each, how would each of them have many moles of  $H^+$  ions and moles of  $OH^-$  ions present at the end point of the reaction?

**PROBLEM** Write the answer on the line to the left. Show all your work in the space provided.

1. A 25-mL solution of 0.100 M  $HCl$  is mixed with a 25-mL solution of sodium hydroxide to determine its molarity and its volume. It takes 25 mL of the base to neutralize the  $HCl$  in the acid.
- a. Write the balanced molecular equation for this neutralization reaction.

# Chapter 15 Acid Base Titration Ph Section 2 Answers

**ML Morrison**



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