

Acids and Bases

Overview of Acids and Bases

Acid-base reactions

Acid-base reactions occur in two proton steps (H⁺)
Base-base reactions occur in two OH⁻ steps



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Acid Strength

Acid strength

Acid strength is determined by the extent to which the acid dissociates in water. The stronger the acid, the more it dissociates.

$$\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$$

Acid strength is determined by the extent to which the acid dissociates in water. The stronger the acid, the more it dissociates.

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Acid Strength and pKa

The Ka value of an acid is related to its acid strength.



The pKa value is related to the Ka value.

To simplify the equation, pKa is used instead of Ka.

$$\text{pKa} = -\log K_a$$

$$\text{pKa}_1 = -\log K_{a1} \quad \text{pKa}_2 = -\log K_{a2}$$

The pKa value is related to the Ka value.

Acid strength is determined by the extent to which the acid dissociates in water.



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base.

The more stable the conjugate base, the stronger the acid.

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Choosing a Base to Deprotonate a Given Compound

Choose a base to deprotonate the following compound.



The pKa of the conjugate acid must be higher than the pKa of the acid.

Choose a base to deprotonate the following compound.



Choosing an Acid to Protonate a Given Compound

Choose an acid to protonate the following compound.

Choose an acid to protonate the following compound.

Choose an acid to protonate the following compound.



The pKa of the conjugate acid must be lower than the pKa of the acid.

Choose an acid to protonate the following compound.

Choose an acid to protonate the following compound.

Choose an acid to protonate the following compound.



Handerson-Hasselbalch Equation

The Henderson-Hasselbalch equation relates the pH of a solution to the pKa of an acid.

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$$\text{pH} = \text{pKa} + \log \frac{[\text{A}^-]}{[\text{HA}]}$$

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Thomas Elliott Taylor



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