

# Annotated periodic table of the elements

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
1 H 1.01 Hydrogen	Main common oxidation states, electronegativity (above-Pauling)																2 He 4.003 Helium	
2 Li 6.94 Lithium	3 Be 9.012 Beryllium																	
3 Na 22.99 Sodium	4 Mg 24.31 Magnesium																	
4 K 39.10 Potassium	5 Ca 40.08 Calcium																	
5 Rb 85.47 Rubidium	6 Sr 87.62 Strontium																	
6 Cs 132.91 Cesium	7 Ba 137.33 Barium																	
7 Fr 223.028 Francium	8 Ra 226.025 Radium																	
Atomic number and symbol above, weight and name below		9 Sc 44.96 Scandium	10 Ti 47.87 Titanium	11 V 50.94 Vanadium	12 Cr 52.00 Chromium	13 Mn 54.94 Manganese	14 Fe 55.85 Iron	15 Co 58.93 Cobalt	16 Ni 58.69 Nickel	17 Cu 63.55 Copper	18 Zn 65.38 Zinc	19 Ga 69.72 Gallium	20 Ge 72.64 Germanium	21 As 74.92 Arsenic	22 Se 78.96 Selenium	23 Br 79.90 Bromine	24 Kr 83.80 Krypton	
		25 Y 88.91 Yttrium	26 Zr 91.22 Zirconium	27 Nb 92.91 Niobium	28 Mo 95.94 Molybdenum	29 Tc 98.91 Technetium	30 Ru 101.07 Ruthenium	31 Rh 101.07 Rhodium	32 Pd 106.42 Palladium	33 Ag 107.87 Silver	34 Cd 112.41 Cadmium	35 In 114.82 Indium	36 Sn 118.71 Tin	37 Sb 121.76 Antimony	38 Te 127.60 Tellurium	39 I 126.91 Iodine	40 Xe 131.29 Xenon	
		57 to 70		71 La 138.91 Lanthanum	72 Hf 178.49 Hafnium	73 Ta 180.95 Tantalum	74 W 183.85 Tungsten	75 Re 186.21 Rhenium	76 Os 190.23 Osmium	77 Ir 192.22 Iridium	78 Pt 195.08 Platinum	79 Au 196.97 Gold	80 Hg 200.59 Mercury	81 Tl 204.38 Thallium	82 Pb 207.2 Lead	83 Bi 208.98 Bismuth	84 Po 209 Polonium	85 At 209 Astatine
		89 to 102		89 Lr 260.11 Lawrencium	90 Rf 261.10 Rutherfordium	91 Db 262.11 Dubnium	92 Sg 263.10 Seaborgium	93 Bh 264.10 Bohrium	94 Hs 265.10 Hassium	95 Mt 266.10 Meitnerium	96 Ds 267.10 Darmstadtium	97 Rg 268.10 Roentgenium	98 Cn 269.10 Copernicium	99 Nh 270.10 Nihonium	100 Fl 271.10 Flerovium	101 Mc 272.10 Moscovium	102 Lv 273.10 Livermorium	103 Ts 274.10 Tennessine

- Group names\*
- 1 Hydrogen & the alkali metals
  - 2 Alkaline earth metals
  - 3 Lanthanoids
  - 4 Actinoids
  - 5 Boron group
  - 6 Carbon group
  - 7 Nitrogen group
  - 8 Oxygen group
  - 9 Halogens
  - 10 Helium group
  - 11 Crystallogens
  - 12 Pnictogens
  - 13 Chalcogens
  - 14 Halogens
  - 15 Helium group

\*Groups 1 to 10 are named after their first members i.e. Group 1 is the Sodium Group

## CLASS COUNTS/PARTS

<b>Metals</b>	Nonmetals	
<b>Active</b>	Pre-transition	Composite
	Alkali earth Alkali	
<b>Transition</b>	Intermediate	
<b>Fluoride</b>	Metalloid	
None (Fr-Fr, Cs-Ac)		
None gas		

89 La 138.91 Lanthanum	90 Ce 140.12 Cerium	91 Pr 140.91 Praseodymium	92 Nd 144.24 Neodymium	93 Pm 144.91 Promethium	94 Sm 150.36 Samarium	95 Eu 151.96 Europium	96 Gd 157.25 Gadolinium	97 Tb 158.93 Terbium	98 Dy 162.50 Dysprosium	99 Ho 164.93 Holmium	100 Er 167.26 Erbium	101 Tm 168.93 Thulium	102 Yb 173.05 Ytterbium
99 Ac 227.03 Actinium	100 Th 232.04 Thorium	101 Pa 231.04 Protactinium	102 U 238.03 Uranium	103 Np 237.05 Neptunium	104 Pu 244.06 Plutonium	105 Am 243.06 Americium	106 Cm 247.07 Curium	107 Bk 247.07 Berkelium	108 Cf 251.08 Californium	109 Es 252.08 Einsteinium	110 Fm 257.09 Fermium	111 Md 258.10 Mendelevium	112 No 259.10 Nobelium

Radioactive  
Approximate data for either a metastable or a ground state is given

**Active metals**

**Pre-transition metals**  
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

**Alkali earths**  
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

**Actinides**  
Soft, dense and reactive metals. Those from Es onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

**Transition metals**  
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

**Noble metals**  
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; inertlike oxides.

**Frontier metals**  
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and ionic species such as zincates, stannates, and bismuthates.

**Metals**

**Nonmetals**

**Metalloids**  
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

**Intermediate nonmetals**  
H and N are colourless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metalloids to near-metallic compounds with metals. Their oxides are acidic, or neutral (H<sub>2</sub>O, CO, NO, N<sub>2</sub>O).

\* as liquid F, the most stable form in ambient conditions

**Group 1**  
The starts as Sc+H, and then bifurcates into a La branch (i.e. La-Ac), and a Lu branch (i.e. Lu-Lr). Trends going down Sc+H, and the Lu branch, resemble those of groups 1-2.

\* Lu branch, more resemble those of groups 4-10

**Composite nonmetals**  
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically composite nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

**Rare gases**  
Colourless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF<sub>6</sub>, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

# Bringing It To The Periodic Table Answers

**Daniel F McAuley**



## **Bringing It To The Periodic Table Answers:**

**Science Tutor: Chemistry, Grades 7 - 8** Raham,2008-08-28 Connect students in grades 7 and up with science using Science Tutor Chemistry This effective 48 page resource provides additional concept reinforcement for students who struggle in chemistry Each lesson in this book contains an Absorb section to instruct and simplify concepts and an Apply section to help students grasp concepts on their own The book covers topics such as matter physical and chemical changes mixtures and solutions the periodic table atomic structure and radioactivity It is great for use in the classroom and at home

Science Tutor: Chemistry, Grades 7 - 12 Gary Raham,2005-01-03 Introduces new chemistry concepts and provides activities so that students can practice and grasp the concepts Key terms are highlighted in the text as well as in a comprehensive glossary Answer keys are included MCAT Critical Analysis and Reasoning Skills Review 2026-2027 Kaplan Test Prep,2025-07-08 Kaplan s MCAT Critical Analysis and Reasoning Skills Review 2026 2027 offers an expert study plan detailed subject review and hundreds of online and in book practice questions all authored by the experts behind Kaplan s score raising MCAT prep course Prepping for the MCAT is a true challenge Kaplan can be your partner along the way offering guidance on where to focus your efforts and how to organize your review This book has been updated to match the AAMC s guidelines precisely no more worrying about whether your MCAT review is comprehensive The Most Practice More than 100 questions in the book and access to even more online more practice than any other MCAT CARS book on the market The Best Practice Comprehensive CARS subject review is written by top rated award winning Kaplan instructors All material is vetted by editors with advanced science degrees and by a medical doctor Online resources including a full length practice test help you practice in the same computer based format you ll see on Test Day Expert Guidance We know the test The Kaplan MCAT team has spent years studying every MCAT related document available Kaplan s expert psychometricians ensure our practice questions and study materials are true to the test IOCL Apprentice (Data Entry Operator) 10 Practice Sets and Solved Papers Book for 2021 Exam with Latest Pattern and Detailed Explanation by Rama Publishers Rama Publishers,2022-02-07 **Academic Computing** ,1987 **Science Examinations ... Reports, Etc** ,1909 **The Chemical News** ,1928 Foundations of College Chemistry, Alternate Morris Hein,Susan Arena,2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **General Chemistry for Colleges** B. Smith Hopkins,1930 **Putting It All Together** Richard H. Bube,1995 A thinking person can hardly avoid developing some kind of pattern within which to express a personal reaction to both scientific and theological inputs Bube works to clarify the

identity of these possible patterns for relating science and the Christian faith to give examples and to provide a balanced critique of each Such an understanding of the issues involved is essential for mature Christians living in a world dominated by the concepts and artifacts of science Contents A Look at the Issues Authentic Science Authentic Christian Theology Pattern 1 Science Has Destroyed Christian Theology Pattern 2 Christian Theology in Spite of Science Pattern 3 Science and Christian Theology are Unrelated Pattern 4 Science Demands Christian Theology Pattern 5 Science Redefines Christian Theology Pattern 6 A New Synthesis of Science and Christian Theology Pattern 7 Christian Theology and Science Complementary Insights Epilogue A Christian Philosophy of Science     **The Pharmaceutical Era** ,1891     **Chemical News and Journal of Industrial Science** ,1928     The Chemical News and Journal of Industrial Science ,1928  
    *Experiments and Exercises in Basic Chemistry* Steven Murov,Brian Stedjee,2000-02-03     *Science & Culture* ,1952  
    Nature Sir Norman Lockyer,1900     **General Chemistry** Hermann Irving Schlesinger,1930     **The National Provisioner** ,1924     **Chemical News and Journal of Physical Science** ,1919     **Educational Times** ,1923

Bringing It To The Periodic Table Answers: Bestsellers in 2023 The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous engrossing novels captivating the hearts of readers worldwide. Lets delve into the realm of top-selling books, exploring the captivating narratives that have enthralled audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover expertly weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed. Bringing It To The Periodic Table Answers : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This intriguing historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Bringing It To The Periodic Table Answers : Delia Owens "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting. These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered. The novel begins with Richard Pape, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is an exceptional and suspenseful novel that will keep you guessing until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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