

Chapter 12

SECTION 12.1

12.1.1 Boyle's Law: Pressure-Volume Relationship for an Ideal Gas

1. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. Show that Boyle's Law is a special case of the following relationship:

$$PV = nRT$$

$$PV = nRT$$

$$PV = nRT$$

$$PV = nRT$$

2. Two sealed flasks, A and B, contain the same amount of gas at equal volume at the same temperature. Flask A is connected to flask B by a tube. What happens to the pressure in the two flasks when the tube is opened? Will the pressure in the two flasks be equal? If not, which flask will have the higher pressure? Will the pressure in the two flasks be equal? If not, which flask will have the higher pressure? Will the pressure in the two flasks be equal? If not, which flask will have the higher pressure?

3. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. Show that Boyle's Law is a special case of the following relationship:

4. A sample of gas is contained in a flask at 10°C. The pressure of the gas is 1.0 atm. The flask is cooled to 0°C. What happens to the pressure of the gas? Will the pressure in the flask be equal to 1.0 atm? If not, what will be the pressure? Will the pressure in the flask be equal to 1.0 atm? If not, what will be the pressure? Will the pressure in the flask be equal to 1.0 atm? If not, what will be the pressure?

5. A sample of gas is contained in a flask at 10°C. The pressure of the gas is 1.0 atm. The flask is cooled to 0°C. What happens to the pressure of the gas? Will the pressure in the flask be equal to 1.0 atm? If not, what will be the pressure? Will the pressure in the flask be equal to 1.0 atm? If not, what will be the pressure? Will the pressure in the flask be equal to 1.0 atm? If not, what will be the pressure?

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